VALENCIACOLLEGE

Experiment 6A: Amount of Active Ingredient in Aspirin

ame:	Date:		
ab Partner:	Section:		
Experimental Data and Calculations			
Table 1. Determination of NaOH Concentration. (Part A du	e next Lab.)		
1) HCl concentration written on carboy			
2) NaOH concentration written on carboy			
3) NaOH Solution A concentration (Theoretical value)			
4) NaOH Solution B concentration (Theoretical value)			
	Trial 1*	Trial 2	Trial 3
5) Volume of HCl sample			
6) Moles HCl			
7) Moles of NaOH			
8) Initial buret reading			
9) Final buret reading			
10) Volume NaOH added			
11) Molar concentration of NaOH Solution B (Experimental Value)			
12) Mean molar concentration of NaOH Solution B			
13) Standard deviation			
14) RSD			
15) Relative percent error (%)			

* Show calculations for Trial 1, standard deviation, RSD and relative % error. (Continue on back of sheet if needed.)

1) Aspirin Brand (and amount of acetylsalicylic acid on label)				
2) Average molar concentration of NaOH Solution B (from Part A)				
	Trial 1	* Tr	ial 2 Ti	ial 3
3) Mass of aspirin tablet				
4) Mass of pulverized aspirin sample				
5) Initial buret reading				
6) Final buret reading				
7) Volume of NaOH Solution B added				
8) Moles of NaOH				
9) Moles of acetylsalicylic acid in sample				
10) Mass of acetylsalicylic acid in sample				
11) Amount of acetylsalicylic acid per tablet				
12) Mean amount of acetylsalicylic acid per tablet				
13) Standard deviation for acetylsalicylic acid per table	t			
14) RSD for acetylsalicylic acid per tablet				
15) Relative percent error (%)				

Data Table 2. Determination of Acetylsalicylic Acid in Aspirin.

*Show all calculations related to Trial 1, standard deviation, RSD, and relative % error.

Report: Include tables 1 and 2, calculations, and written conclusion.